Daily Tutorial Sheet-6	Level-2

76.(A) In (b, c, d) carbon show + 4 oxidation state while in (A) carbon show – 4 oxidation state.

77.(A) $\overset{x}{C}H_2Cl_2 \qquad x + 2 - 2 = 0 \Rightarrow x = 0$

78.(B) $\overset{*}{S}O_2 = +4 \;\; ; \;\; H_2\overset{*}{S}O_4 = +6 \;\; ; \;\; Na_2\overset{*}{S}_2O_3 = +2 \;\; ; \;\; Na_2\overset{*}{S}_4O_6 = +\frac{5}{2}.$

79.(B) Oxidation number of C in

$HCHO = 0 \;\; ; \;\; CHCl_3 = +2 \;\; ; \;\; CH_3OH = -2 \;\; ; \;\; C_{12}H_{22}O_{11} = 0$

80.(D) Fluorine always shows – 1 oxidation state.

81.(B) In all alkali and alkaline earth metal hydride hydrogen always shows – 1 oxidation state.

82.(C) Oxygen has 6 electrons in the outer most shell and shows common oxidation state – 2.

83.(C) $\overset{\text{Oxidation}}{\overset{0}{H_2} + \overset{0}{Br_2} \longrightarrow 2\overset{+}{H} - \overset{-}{Br}}$ (Reduction)

84.(C) Al_2O_3 cannot act as both oxidising and reducing agent.

85.(AB) In H_2S sulphur shows –2 oxidation state and in SO_2 shows +4 oxidation state. Hence SO_2 shows both oxidising and reducing properties.